

## 14 The Behavior Of Gases Chapter Quiz

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### 14 The Behavior Of Gases

the rates of effusion of gases (at the same temperature and pressure) are inversely proportional to the square roots of their molar masses Diffusion the tendency of molecules to move toward areas of lower concentration until the concentration is uniform throughout

### Chemistry: Chapter 14: The Behavior of Gases Flashcards ...

at a constant volume and temperature, the total pressure exerted by a mixture of gases is equal to the sum of the partial pressures of the components of gases diffusion the tendency of molecules to move toward areas of lower concentration until the concentration is uniform throughout

### Chapter 14: Behavior of Gases Flashcards | Quizlet

Chapter 14 The Behavior of Gases 147 SECTION 14.1 PROPERTIES OF GASES (pages 413–417) This section uses kinetic theory to explain the properties of gases. This section also explains how gas pressure is affected by the amount of gas, its volume, and its temperature. Compressibility (pages 413–414) 1. Look at Figure 14.1 on page 413.

### SECTION 14.1 PROPERTIES OF GASES (pages 413–417)

At a constant volume and temperature, the total pressure exerted by a mixture of gases is equal to the sum of the partial pressures of the component gases. Diffusion The tendency of molecules to move toward areas of lower concentration until the concentration is uniform throughout

### Chapter 14: Behavior of Gases Flashcards | Quizlet

Chapter 14 - The Behavior of Gases - 14.1 Properties of Gases - 14.1 Lesson Check - Page 454: 6 Answer If the temperature is constant, quadrupling the volume would cause the pressure of an enclosed gas to be reduced to one quarter of its original value.

### Chapter 14 - The Behavior of Gases - 14.1 Properties of ...

Gases are all around us and they are all unique in their own way be it in behavior or location they are found. In chapter fourteen we got to understand the behavior of various gasses and their qualities. The quiz below is designed to test how well you understood that. Give it a try!

### The Behavior Of Gases Chapter 14 - ProProfs Quiz

Gases are compressible because most of the volume of a gas is composed of the large amounts of empty space between the gas particles. At room temperature and standard pressure, the average distance between gas molecules is about ten times the diameter of the molecules themselves.

### 14.1: Compressibility - Chemistry LibreTexts

None of the gases that exist in nature, follow the gas laws for all values of temperature and pressure. So we see that the behavior of gases that exist or the "real gases" differs from the behavior of the ideal gases. These gases deviate from ideal gas laws because: Real gas molecules attract one another.

### **Behavior of Gases: Laws, Ideal & Real Gas, Concepts ...**

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Gases are easy to compress because of the empty space between the particles in a gas. Why is a gas easy to compress? the amount of gas, volume and temperature are all factors that affect gas pressure.

### **Chapter 14 Gas Laws Review Flashcards | Quizlet**

The Chapter 14 Behavior of Gases Section 14.1 Properties of Gases \ Converting Pressure 1 atmosphere = 760 mmHg 1 atmosphere = 101 325 Pa 1 atmosphere = 101.325 kPa ... - PowerPoint PPT presentation

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Chapter 14 The Behavior of Gases Chemistry Honors \* \* Avagadro s Hypothesis Equal volumes of gas (at same T and P) contain the same amount of particles \* 1 mole = 6 ... - PowerPoint PPT presentation

### **PPT - Chapter 14 The Behavior of Gases PowerPoint ...**

KINETIC THEORY OF GASES | Laws governing kinetic behavior of gases. Boyle's law gives a relationship between pressure and volume of a gas at constant temperature.It states that the volume of a given mass of a gas.....

### **KINETIC THEORY OF GASES | Laws governing kinetic behavior ...**

Title: Chapter 14 The Behavior of Gases 1 Chapter 14 The Behavior of Gases. Pre-AP Chemistry ; Charles Page High School ; Stephen L. Cotton; 2 Section 14.1 The Properties of Gases. OBJECTIVES ; Explain why gases are easier to compress than solids or liquids are. 3 Section 14.1 The Properties of Gases. OBJECTIVES ; Describe the three factors ...

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### **14: The Behavior of Gases - Chemistry LibreTexts**

Title: Chapter 14 The Behavior of Gases 1 Chapter 14 The Behavior of Gases 2 Section 14.2The Gas Laws. \ 3 Converting Pressure . 1 atmosphere 760 mmHg ; 1 atmosphere 101 325 Pa ; 1 atmosphere 101.325 kPa

### **PPT - Chapter 14 The Behavior of Gases PowerPoint ...**

Specific laws that express the behavior of gases at various temperature and pressure are very important in petroleum technology. The gases are divided into ideal (perfect) gas and real ( non-ideal) gas. The properties of hydrocarbon gases are relatively simple since the parameters of pressure, Volume and temperature (PVT) can be related by a single equation.

### **Behavior of Gases - LinkedIn SlideShare**

Lab 14: Studying Boyle's Law 3. Boyle's law describes the behavior of ideal gases. When

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condensed, ideal gases have insignificant molecular volumes when compared to the volume to which the gas is confined. Also, no attractions exist between ideal gas molecules.

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